AQA

AS-LEVEL
Chemistry

AS Paper 1
Mark scheme

7404/1
June 2017

Version: 1.0 Final

Mark schemes are prepared by the Lead Assessment Writer and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all associates participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the students' responses to questions and that every associate understands and applies it in the same correct way. As preparation for standardisation each associate analyses a number of students' scripts. Alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, associates encounter unusual answers which have not been raised they are required to refer these to the Lead Assessment Writer.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of students' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

Further copies of this mark scheme are available from aqa.org.uk

AS-Level Chemistry
Mark Scheme Instructions for Examiners

1. General

The mark scheme for each question shows:

- the marks available for each part of the question
- the total marks available for the question
- the typical answer or answers which are expected
- extra information to help the examiner make his or her judgement and help to delineate what is acceptable or not worthy of credit or, in discursive answers, to give an overview of the area in which a mark or marks may be awarded.

The extra information in the 'Comments' column is aligned to the appropriate answer in the left-hand part of the mark scheme and should only be applied to that item in the mark scheme.

You should mark according to the contents of the mark scheme. If you are in any doubt about applying the mark scheme to a particular response, consult with your Team Leader.

At the beginning of a part of a question a reminder may be given, for example: where consequential marking needs to be considered in a calculation; or the answer may be on the diagram or at a different place on the script.

In general the right-hand side of the mark scheme is there to provide those extra details which confuse the main part of the mark scheme yet may be helpful in ensuring that marking is straightforward and consistent.

The use of M1, M2, M3 etc in the right-hand column refers to the marking points in the order in which they appear in the mark scheme. So, M1 refers to the first marking point, M2 the second marking point etc.

2. Emboldening

2.1 In a list of acceptable answers where more than one mark is available 'any **two** from' is used, with the number of marks emboldened. Each of the following bullet points is a potential mark.

2.2 A bold **and** is used to indicate that both parts of the answer are required to award the mark.

2.3 Alternative answers acceptable for a mark are indicated by the use of **OR**. Different terms in the mark scheme are shown by a / ; eg allow smooth / free movement.

Marking points

3.1 Marking of lists

This applies to questions requiring a set number of responses, but for which students have provided extra responses. The general principle to be followed in such a situation is that 'right + wrong = wrong'.

Each error / contradiction negates each correct response. So, if the number of error / contradictions equals or exceeds the number of marks available for the question, no marks can be awarded.

However, responses considered to be neutral (often prefaced by 'Ignore' in the mark scheme) are not penalised.

For example, in a question requiring 2 answers for 2 marks:

Correct answers	Incorrect answers (ie incorrect rather than neutral)	Mark (2)	Comment
1	0	1	
1	1	1	They have not exceeded the maximum number of responses so there is no penalty.
1	2	0	They have exceeded the maximum number of responses so the extra incorrect response cancels the correct one.
2	0	2	
2	1	1	
2	2	0	
3	0	2	The maximum mark is 2
3	1	1	The incorrect response cancels out one of the two correct responses that gained credit.
3	2	0	Two incorrect responses cancel out the two marks gained.
3	3	0	

3.2 Marking procedure for calculations

Full marks should be awarded for a correct numerical answer, without any working shown, unless the question states 'Show your working' or 'justify your answer'. In this case, the mark scheme will clearly indicate what is required to gain full credit.

If an answer to a calculation is incorrect and working is shown, process mark(s) can usually be gained by correct substitution / working and this is shown in the 'Comments' column or by each stage of a longer calculation.

3.3 Errors carried forward, consequential marking and arithmetic errors

Allowances for errors carried forward are most likely to be restricted to calculation questions and should be shown by the abbreviation ECF or consequential in the marking scheme.

An arithmetic error should be penalised for one mark only unless otherwise amplified in the marking scheme. Arithmetic errors may arise from a slip in a calculation or from an incorrect transfer of a numerical value from data given in a question.

3.4 Extended responses

Where a mark scheme includes linkage words (such as 'therefore', 'so', 'because' etc), these are optional. However, a student's marks for the question may be limited if they do not demonstrate the ability to construct and develop a sustained line of reasoning which is coherent, relevant, substantiated and logically structured. In particular answers in the form of bullet pointed lists may not be awarded full marks if there is no indication of logical flow between each point or if points are in an illogical order.

3.5 Equations

In questions requiring students to write equations, state symbols are generally ignored unless otherwise stated in the 'Comments' column.

Examiners should also credit correct equations using multiples and fractions unless otherwise stated in the 'Comments' column.

3.6 Oxidation states

In general, the sign for an oxidation state will be assumed to be positive unless specifically shown to be negative.

3.7 Interpretation of 'it'

Answers using the word 'it' should be given credit only if it is clear that the 'it' refers to the correct subject.

3.8 Phonetic spelling

The phonetic spelling of correct scientific terminology should be credited **unless** there is a possible confusion with another technical term or if the question requires correct IUPAC nomenclature.

3.9 Brackets

(…..) are used to indicate information which is not essential for the mark to be awarded but is included to help the examiner identify the sense of the answer required.

3.10 Ignore / Insufficient / Do <u>not</u> allow

Ignore or insufficient is used when the information given is irrelevant to the question or not enough to gain the marking point. Any further correct amplification could gain the marking point.

Do **not** allow means that this is a wrong answer which, even if the correct answer is given, will still mean that the mark is not awarded.

3.11 Marking crossed out work

Crossed out work that **has not been** replaced should be marked as if it were not crossed out, if possible. Where crossed out work **has been** replaced, mark the replacement work and not the crossed out work.

3.12 Reagents

The command word "Identify", allows the student to choose to use **either** the name or the formula of a reagent in their answer. In some circumstances, the list principle may apply when both the name and the formula are used. Specific details will be given in mark schemes.

The guiding principle is that a reagent is a chemical which can be taken out of a bottle or container. Failure to identify complete reagents **will be penalised**, but follow-on marks (e.g. for a subsequent equation or observation) can be scored from an incorrect attempt (possibly an incomplete reagent) at the correct reagent. Specific details will be given in mark schemes.

For example, **no credit** would be given for

- the cyanide ion or CN^- when the reagent should be potassium cyanide or KCN;
- the hydroxide ion or OH^- when the reagent should be sodium hydroxide or NaOH;
- the $Ag(NH_3)_2^+$ ion when the reagent should be Tollens' reagent (or ammoniacal silver nitrate). In this example, no credit is given for the ion, but credit could be given for a correct observation following on from the use of the ion. Specific details will be given in mark schemes.

In the event that a student provides, for example, **both** KCN and cyanide ion, it would be usual to ignore the reference to the cyanide ion (because this is not contradictory) and credit the KCN. Specific details will be given in mark schemes.

3.13 Organic structures

Where students are asked to draw organic structures, unless a specific type is required in the question and stated in the mark scheme, these may be given as displayed, structural or skeletal formulas or a combination of all three as long as the result unambiguous.

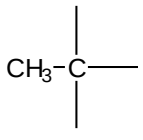
In general

- Displayed formulae must show all of the bonds and all of the atoms in the molecule, but need not show correct bond angles.
- Skeletal formulae must show carbon atoms by an angle or suitable intersection in the skeleton chain. Functional groups must be shown and it is essential that all atoms other than C atoms are shown in these (except H atoms in the functional groups of aldehydes, secondary amines and N-substituted amides which do not need to be shown).
- Structures must not be ambiguous, e.g. 1-bromopropane should be shown as $CH_3CH_2CH_2Br$ and not as the molecular formula C_3H_7Br which could also represent the isomeric 2-bromopropane.
- Bonds should be drawn correctly between the relevant atoms. This principle applies in all cases where the attached functional group contains a carbon atom, e.g nitrile,

carboxylic acid, aldehyde and acid chloride. The carbon-carbon bond should be clearly shown. Wrongly bonded atoms will be penalised **on every occasion**. (see the examples below)

- The same principle should also be applied to the structure of alcohols. For example, if students show the alcohol functional group as C ─ HO, they should be penalised **on every occasion**.
- Latitude should be given to the representation of C ─ C bonds in alkyl groups, given that CH_3─ is considered to be interchangeable with H_3C─ even though the latter would be preferred.
- Similar latitude should be given to the representation of amines where NH_2─ C will be allowed, although H_2N─ C would be preferred.
- Poor presentation of vertical C ─ CH_3 bonds or vertical C ─ NH_2 bonds should **not** be penalised. For other functional groups, such as ─ OH and ─ CN, the limit of tolerance is the half-way position between the vertical bond and the relevant atoms in the attached group.

By way of illustration, the following would apply.



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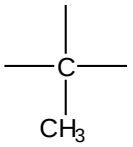

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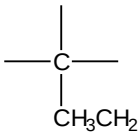

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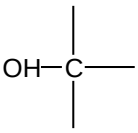

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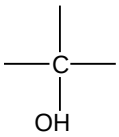

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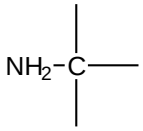

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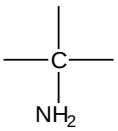

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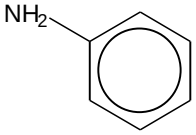

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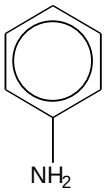

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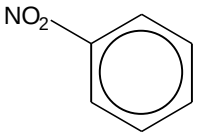

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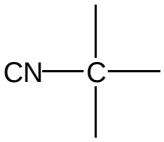

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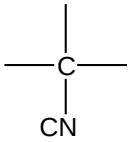

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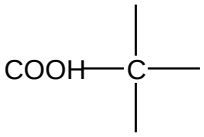

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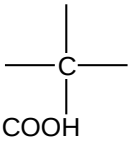

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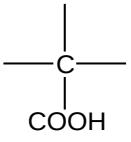

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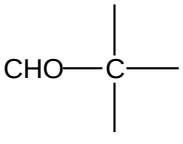

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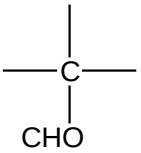

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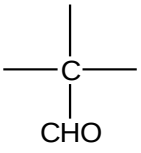

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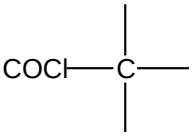

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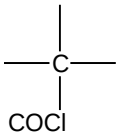

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- Representation of CH_2 by $C{-}H_2$ will be penalised
- Some examples are given here of **structures** for specific compounds that should **not** gain credit (but, exceptions may be made in the context of balancing equations)

CH_3COH	for	ethanal
CH_3CH_2HO	for	ethanol
$OHCH_2CH_3$	for	ethanol
C_2H_6O	for	ethanol
CH_2CH_2	for	ethene
$CH_2.CH_2$	for	ethene
$CH_2:CH_2$	for	ethane

- Each of the following **should gain credit** as alternatives to correct representations of the structures.

$CH_2 = CH_2$	for	ethene, $H_2C=CH_2$
$CH_3CHOHCH_3$	for	propan-2-ol, $CH_3CH(OH)CH_3$

- In most cases, the use of "sticks" to represent C ─ H bonds in a structure should **not** be penalised. The exceptions to this when "sticks" will be penalised include
  - structures in mechanisms where the C ─ H bond is essential (e.g. elimination reactions in halogenoalkanes and alcohols)
  - when a displayed formula is required
  - when a skeletal structure is required or has been drawn by the candidate

3.14 Organic names

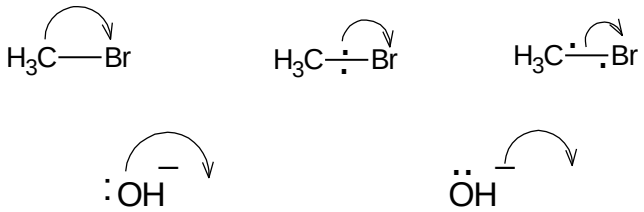
As a general principle, non-IUPAC names or incorrect spelling or incomplete names should **not** gain credit. Some illustrations are given here.

but-2-ol	should be **butan-2-ol**
2-hydroxybutane	should be **butan-2-ol**
butane-2-ol	should be **butan-2-ol**
2-butanol	should be **butan-2-ol**
ethan-1,2-diol	should be **ethane-1,2-diol**
2-methpropan-2-ol	should be **2-methylpropan-2-ol**
2-methylbutan-3-ol	should be **3-methylbutan-2-ol**
3-methylpentan	should be **3-methylpentane**
3-mythylpentane	should be **3-methylpentane**
3-methypentane	should be **3-methylpentane**
propanitrile	should be **propanenitrile**
aminethane	should be **ethylamine** (although aminoethane can gain credit)
2-methyl-3-bromobutane	should be **2-bromo-3-methylbutane**
3-bromo-2-methylbutane	should be **2-bromo-3-methylbutane**
3-methyl-2-bromobutane	should be **2-bromo-3-methylbutane**
2-methylbut-3-ene	should be **3-methylbut-1-ene**
difluorodichloromethane	should be **dichlorodifluoromethane**

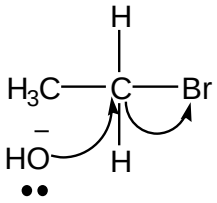
3.15 Organic reaction mechanisms

Curly arrows should originate either from a lone pair of electrons or from a bond.

The following representations should not gain credit **and will be penalised each time** within a clip.



For example, the following would score zero marks



When the curly arrow is showing the formation of a bond to an atom, the arrow can go directly to the relevant atom, alongside the relevant atom or **more than half-way** towards the relevant atom.

In free-radical substitution

- the absence of a radical dot should be penalised **once only** within a clip.
- the use of half-headed arrows is not required, but the use of double-headed arrows or the incorrect use of half-headed arrows in free-radical mechanisms should be penalised **once only** within a clip

The correct use of skeletal formulae in mechanisms is acceptable, but where a C-H bond breaks both the bond and the H must be drawn to gain credit.

Question	Marking Guidance	Mark	Additional Comments/Guidance
01.1	Cl^- $1s^2 2s^2 2p^6 3s^2 3p^6$ Fe^{2+} $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6$	1 1	If [Ne] or [Ar] used then Max 1if both correct Ignore $4s^0$ Allow subscripts
01.2	Mn^{2+} (g) → Mn^{3+} (g) + e^-	1	States symbols are required Allow Mn^{2+} (g) - e^- → Mn^{3+} (g) Negative charge needed on electron
01.3	Al (Outer) electron in (3)p sublevel/orbital Higher in energy/further from the nucleus so easier to remove OWTTE	1 1 1	Mg then CE=0 Not just level or shell Both required for M3 Ignore shielding
01.4	$^{58}Ni^+$ A_r= [(58x61.0)+(60x29.1)+(61x9.9)]/100 A_r= 58.9 must be to 1DP	1 1 1	M1 needs mass and charge – allow subscripts

Question	Marking Guidance	Mark	Additional Comments/Guidance
02.1	$2Fe(s) + {}^{3}/_{2}O_2(g) \rightarrow Fe_2O_3(s)$ ONLY	1	Don't allow multiples. States must be shown
02.2	M1 Correct cycle or equation M2 $(3x\Delta_f HCO_2) = -19 + (-822) + 3(-111) - 0$ $(3x\Delta_f HCO_2) = -1174$ M3 $\Delta_f HCO_2 = -391$ kJmol^{-1}	1 1 1	If M1 and M2 not awarded then M3 can be awarded for their M2 divided by 3 -317 for 1 mark +391 for 1 mark Allow 2 sig fig or more
02.3	M1 Correct Hess's law cycle or equation M2 $(6(N\text{-}H)) = 944 + 3(+436) + 92$ $(6(N\text{-}H)) = 2344$ M3 N-H = (+)391kJmol^{-1}	1 1 1	If M1 and M2 not awarded then M3 can be awarded for their M2 divided by 6 -391 for 1 mark Allow 2 sig fig or more
02.4	Data book value derived from (a number of) different compounds (not just different NH_3 molecules)	1	

Question	Marking Guidance	Mark	Additional Comments/Guidance
03.1	M1 Amount $ZnSO_4$ = 1.0 x $^{50}/_{1000}$ mol or Amount $ZnSO_4$ =0.050 mol	1	Mark M1 and M2 independently
	M2 Amount Mg = $^{2.08}/_{24.3}$ mol or Amount Mg= 0.0856 mol	1	
	(Hence Mg in excess)		
	M3 $Q=mc\Delta T$	1	M3 could be scored in M4
	M4 Q=50.0 x 4.18 x 37.3 or Q=7795.7J	1	If an error in M4, lose M4 and M5 and only award M6 for correct use of their incorrect M4 and division by their correct limiting reagent
	M5 (Energy released per mole) = $^{7.796}/_{0.05}$kJmol^{-1} or $^{7796}/_{0.05}$ Jmol^{-1}	1	M5 division by their limiting reagent
	M6 ΔH= − 156 kJmol^{-1}	1	
03.2	Heat loss (from the apparatus would mean the experimental value is smaller/lower/less exothermic than the data source)	1	

Question	Marking Guidance		Mark	Comments
03.3	Marks awarded for this answer will be determined by the as well as the standard of the scientific response. Examiners should apply a 'best-fit' approach to the marking.		6	**Indicative Chemistry Content** **Stage 1 Improved insulation** 1a Insulate the beaker or use a polystyrene cup or a lid 1b To reduce heat loss **Stage 2 Improved temperature recording** 2a Record the temperature for a suitable time before adding the metal 2b To establish an accurate initial temperature OR 2c Record temperature values at regular time intervals 2d To plot the temperature results against time on a graph **Stage 3 Improved analysis of results** 3a Extrapolate the cooling back to the point of addition 3b To establish a (theoretical) maximum temperature OR temperature change (e.g. at the 4th minute) OR adjust for the cooling /apply a cooling correction 3a and 3b could be seen on an extrapolated sketch graph (Note- IGNORE use of measuring equipment with greater precision)
	Level 3 (5—6 marks)	**Covers 3 Stages with matching justifications** Answer is full and detailed and is supported by an appropriate range of relevant points such as those given below: - argument is well structured with minimum repetition or irrelevant points - accurate and clear expression of ideas with only minor errors in the use of technical terms, spelling and punctuation and grammar		
	Level 2 (3—4 marks)	**Covers 2 Stages with matching justification. OR covers 3 Stages with incomplete justification** Answer has some omissions but is generally supported by some of the relevant points below: - the argument shows some attempt at structure - the ideas are expressed with reasonable clarity but with a few errors in the use of technical terms, spelling, punctuation and grammar		
	Level 1 (1—2 marks)	**Covers 1 Stage with matching justification. OR covers 2 Stages with incomplete justification** Answer is largely incomplete. It may contain valid points which are not clearly linked to an argument structure. Unstructured answer. Errors in the use of technical terms, spelling, punctuation and grammar or lack of fluency		

	Level 0 **0 marks**	**Insufficient correct chemistry**		

Question	Marking Guidance	Mark	Additional Comments/Guidance
04.1	mol R=2x	1	
04.2	$3.6 = \frac{(2x)^2}{(1-x)^2}$ $\sqrt{3.6} = \frac{2x}{1-x}$ (only positive root to be used) $\sqrt{3.6} - \sqrt{3.6}\,x = 2x$ $1.9 = 3.9x$ $X=0.49$ $[R] = 0.97 mol dm^{-3}$ (allow range 0.97-.098)	1 1 1	M1 can be awarded for the insertion of their answer from 04.1 correctly M2 can be awarded if their expression is expanded M3 solve for x from their expression in M1 and use it to calculate [R]

Question	Marking Guidance	Mark	Additional Comments/Guidance
05.1	Power of an atom to attract a pair of electrons in a covalent bond.	1	Allow power of an atom to attract a bonding/shared pair of electrons Allow power of an atom to withdraw electron density from a covalent bond Not lone pair Not Element
05.2	<u>Difference in electronegativity</u> leads to bond polarity (dipoles don't cancel the molecule has an overall permanent dipole) and there is an attraction between ∂+ on one molecule and ∂- on another	1 1	If chloride (ions) mentioned then CE=0 partial charges should be correct if shown and can score M2 from diagram

05.3	SiH_4	**Tetrahedral**		1 shape & no tick	If shapes are drawn rather than named then penalise first mark gained
	PH_3	**Pyramidal (trigonal) Allow tetrahedral**	√	1 shape & tick	
	$BeCl_2$	**Linear**		1 shape & no tick	
	CH_3Cl	**(Distorted)Tetrahedral**	√	1 shape & tick	

Question	Marking Guidance	Mark	Additional Comments/Guidance
06.1	$4CuFeS_2 + 9^1/_2\ O_2 + 4SiO_2 \rightarrow Cu_2S + Cu_2O + 7SO_2 + 4FeSiO_3$ $Cu_2S + 2Cu_2O \rightarrow 6Cu + SO_2$	1 1	Allow multiples
06.2	ANY TWO - Prevents acid rain (which damages buidlings/ecology) - Toxic OR causes breathing problems - Reduces waste product OR makes use of the waste OR improves atom economy OR Reduces need for sulfur mining OR used to produce sulphuric acid OR any named products	1 1	

Question	Marking Guidance	Mark	Additional Comments/Guidance
06.3	M1,M2,M3 are process marks M1 Mol Cu = $\frac{4050 \times 1000}{63.5}$ (= 63780) M2 Mass $CuFeS_2$ = (63780) x 183.5 (= 1.17×10^7g) M3 Mass ore = (1.17×10^7) x $^{100}/_{1.25}$ M4 Mass ore = 936 tonnes (Allow 936 -937)	 1 1 1 1	Alternative method M1 % of Cu in $CuFeS_2$ =(63.5/183.5)x100 = 34.6% M2 % of Cu in the rock=(34.6/100) x 1.25 = 0.4325% M3 mass of rock = 4050 x 100/0.4325 = 936416kg M4 mass of rock in tonnes= 936 tonnes Notes M1 A_r Cu must be used M2 M_r $CuFeS_2$ to have been used M3 Grossing up for the mass of rock M4 Final answer correct in tonnes
06.4	% atom economy = $^{(2x63.5)}/_{171}$ x100 =74.3% must be 3sf	1 1	

Question	Marking Guidance	Mark	Additional Comments/Guidance
07.1	Ba^{2+} OR Sr^{2+} SO_4^{2-} (aq) + Ba^{2+} (aq) → $BaSO_4$ (s)	1 1	Award M1 if barium named in M1 then used Ba^{2+} in the equation
07.2	OH^- Mg^{2+} (aq) + $2OH^-$ (aq) → $Mg(OH)_2$ (s)	1 1	Award M1 if hydroxide named in M1 then used OH^- in the equation

Question	Marking Guidance	Mark	Additional Comments/Guidance
08.1	NO_2^- +3 or III or 3 or 3+ NO +2 or II or 2 or 2+	1 1	
08.2	$NO_2^- + e^- + 2H^+ \rightarrow NO + H_2O$ (OR double)	1	
08.3	$2I^- \rightarrow I_2 + 2e^-$ (OR half)	1	
08.4	$2NO_2^- + 2I^- + 4H^+ \rightarrow I_2 + 2NO + 2H_2O$	1	
08.5	Oxidising agent	1	Allow to accept/gain electrons Allow Oxidant Do not allow accept/ gain pairs of electrons Do not allow Oxidise

08.6	Mol ClO_3^- = $0.02 \times {}^{27.4}/_{1000} = 5.48 \times 10^{-4}$	1	
	Mol NO_2^- = ${}^{5}/_{2}\left(0.02 \times {}^{27.4}/_{1000}\right) = 1.37 \times 10^{-3}$	1	
	$[NO_2^-]$ = mol NO_2^- / $({}^{25}/_{1000})$ $[NaNO_2]$ = 0.0548 moldm^{-3}	1	
	Conc $NaNO_2$ = (0.0548) x 69.0 = 3.78 gdm^{-3}	1	Minimum 2 sf

9	A
10	C
11	C
12	A
13	B
14	B
15	D
16	B
17	B
18	C

19	D
20	B
21	D
22	C
23	A